Interaction of Oxides of Tin with Isopropyl Alcohol

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Interaction of isopropyl alcohol with tin and its oxides was studied at various temperatures in the vapor phase. Frost's equations were found to apply to some of the experimental results. Dependence between conversion and space velocity, and calculated activation energies over various oxides of tin are given. The magnitudes obtained show that the reaction occurs within the kinetic region.

Rates of Combustion of Carbonaceous Deposits in Regeneration of a Butane Dehydrogenation Catalyst

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Combustion of carbonaceous deposits is second order relative to "carbon," and first order with respect to oxygen between 300° and 350°, and activation energy is 31.1 kcal/mole. The rate constant is a function of composition of the carbonaceous deposit, and increases with hydrogen content. Above 400° the rate becomes increasingly more diffusion limited.

Study of Conversion of 2,2,4-Trimethyl Pentane and of n-Octane Over a Platinum Catalyst

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Conversion of 2,2,4-trimethyl pentane and of *n*-octane was compared over a commercial platinum catalyst. The dehydrocyclization mechanism for molecules lacking six-carbon chains is indicated. Highly branched hydrocarbons decompose more readily, form more gaseous decomposition products and aromatics with carbon number smaller than that of the feed.